**Balancing Chemical Equations Practice**

1. \_\_\_\_\_N2 + \_\_\_\_\_O2 → \_\_\_\_\_ N2O
2. \_\_\_\_\_HgO → \_\_\_\_\_Hg + \_\_\_\_\_ O2
3. \_\_\_\_\_ SiCl4 + \_\_\_\_\_H2O → \_\_\_\_\_H4SiO4 + \_\_\_\_\_ HCl
4. \_\_\_\_\_ Na + \_\_\_\_\_ H2O → \_\_\_\_\_NaOH + \_\_\_\_\_ H2
5. \_\_\_\_\_ CO2 + \_\_\_\_\_ NH3 → \_\_\_\_\_CO(NH2)2 + \_\_\_\_\_H2O
6. \_\_\_\_\_Al(OH)3 + \_\_\_\_\_ H2SO4 → \_\_\_\_\_ Al2(SO4)3 + \_\_\_\_\_ H2O
7. \_\_\_\_\_ Fe2(SO4)3 + \_\_\_\_\_ KOH → \_\_\_\_\_K2SO4 + \_\_\_\_\_Fe(OH)3
8. \_\_\_\_\_ Na2CO3 + \_\_\_\_\_ HCl → \_\_\_\_\_NaCl + \_\_\_\_\_ H2O + \_\_\_\_\_CO2
9. \_\_\_\_\_FeS2 + \_\_\_\_\_ O2 → \_\_\_\_\_ Fe2O3 + \_\_\_\_\_SO2
10. \_\_\_\_\_C7H6O2 + \_\_\_\_\_O2 → \_\_\_\_\_ CO2 + \_\_\_\_\_ H2O
11. Dicarbon dihydride + Oxygen → Carbon dioxide + Water
12. Potassium oxide + Water → Potassium Hydroxide
13. Aluminum + Oxygen → Aluminum oxide
14. Iron (III) chloride + Ammonium hydroxide → Iron (III) hydroxide + Ammonium
15. Calcium chlorate → Calcium chloride + Oxygen

**Want more of a challenge? Look on the back!**

**BONUS CHALLENGE QUESTIONS:**

1. \_\_\_\_\_ S + \_\_\_\_\_ HNO3 → \_\_\_\_\_ H2SO4  + \_\_\_\_\_ NO2 + \_\_\_\_\_ H2O
2. \_\_\_\_\_ Cu + \_\_\_\_\_ HNO3 → \_\_\_\_\_ Cu(NO3)2 + \_\_\_\_\_NO + \_\_\_\_\_ H2O
3. \_\_\_\_\_ NaBr + \_\_\_\_\_ NaBrO3 + \_\_\_\_\_ H2SO4 → \_\_\_\_\_ Br2 + \_\_\_\_\_ Na2SO4 + \_\_\_\_\_ H2O
4. \_\_\_\_\_ FeCr2O4 + \_\_\_\_\_ Na2CO3 + \_\_\_\_\_ O2 → \_\_\_\_\_ Na2CrO4 + \_\_\_\_\_ Fe2O3 + \_\_\_\_\_ CO2
5. \_\_\_\_\_ XeF4 + \_\_\_\_\_ H2O → \_\_\_\_\_ XeO3  + \_\_\_\_\_ Xe + \_\_\_\_\_ HF + \_\_\_\_\_ O2

*\*Note: Although Noble gases such as Xenon are normally unreactive, they can sometimes form compounds. Xenon tetrafluoride was the first discovered binary compound of a Noble gas and it can be formed under very high temperatures.*