**Pre-Lab:** What would you look for in each type of reaction to help you determine if that type of reaction was occurring? Use any notes you have or pages 366-367 in the White Chemistry Textbook.

Combustion:

Synthesis:

Decomposition:

Single-Replacement:

Double-Replacement:

**Write Your Hypothesis:**

**Data Table**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **Reaction**  | **Observation of Reaction** | **Endothermic or Exothermic** | **Reaction Type** | **Reactants** | **Products** | **Balanced Chemical Equation** |
| **#1** | Magnesium strip burns, gives off very bright light, turns to white powder |  |  | Magnesium and Oxygen | Magnesium Oxide |  |
| **#2** | **Before adding KI:**Nothing**After KI:**Colored Foam is produced, give off heat |  |  | Hydrogen Peroxide (H2O2) | Water and Oxygen (O2) |  |
| **#3** | Brown flakes start to form on aluminum, some heat felt |  |  | Aluminum and Copper (II) Cloride (CuCl2) | Copper and Aluminum Chloride (AlCl3) |  |
| **#4** | Solution turns purple |  |  | Cobalt (II) Chloride (CuCl2) and Sodium Phosphate (Na3PO4) | Sodium Chloride (NaCl) and Cobalt (II) Phosphate (Co3(PO4)2) |  |