

Identifying atoms in a formula

Complete the chart below. Remember that subscripts outside of a parenthesis get distributed to everything inside the parenthesis and that coefficients get distributed to everything in the formula.

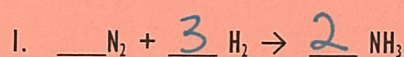
Key

Compound Name	Formula	Number of Elements	Number of atoms of each element
1. Calcium nitrate	$\text{Ca}(\text{NO}_3)_2$	3	Ca=1 N=2 O=6
2. Dihydrogen dioxide	$3\text{H}_2\text{O}_2$	2	H=6 O=6
3. Tin (II) sulfite	2SnSO_3	3	Sn=2 S=2 O=6
4. Ammonium oxalate	$3(\text{NH}_4)_2\text{C}_2\text{O}_4$	4	N=6 C=6 H=24 O=12
5. Pentacarbon Decahydride	$4\text{C}_5\text{H}_{10}$	2	C=20 H=40

Balancing Chemical Equations

Practice Problems

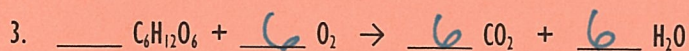
Instructions: Complete the following reactions, balance them properly and identify the type of reaction occurring.



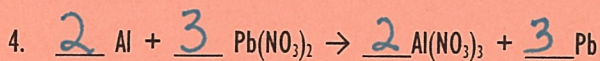
Reaction Type: Combination/Synthesis



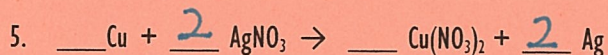
Reaction Type: Decomposition



Reaction Type: Combustion



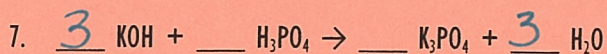
Reaction Type: Single Replacement



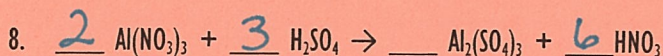
Reaction Type: Single Replacement



Reaction Type: Double Replacement



Reaction Type: Double Replacement



Reaction Type: Double Replacement



Reaction Type: Double Replacement



Reaction Type: Double Replacement